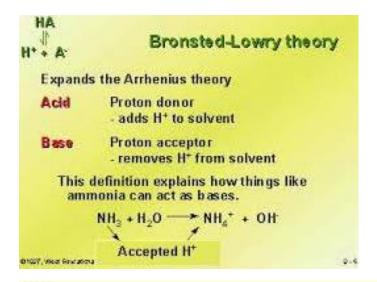
Bronsted-Lowry Theory

Introduction: Bronsted-Lowry theory states that an acid is a proton donor and a base is a proton acceptor.



Relative Strength of Acids & Bases

0 0

Strong acids are better H⁺ donors than weak acids.

Strong bases are better H⁺ acceptors than weak bases.

- Strong acids have weak conjugate bases.
- Weak acids have strong conjugate bases.

Strong acid + H₂O ----- H₃O+ + weak conjugate base

Fully ionized, reverse reaction essentially does not occur.

The conjugate base is weak.

Weak acid + H₂O ← → H₃O⁺ + strong conjugate base

Weakly ionized, reverse reaction readily occurs.

The conjugate base is strong.

https://www.youtube.com/watch?v=ZiokqP0aZ1E&feature=emb_logo

https://prezi.com/f7yktaj5axn7/acids-bases/

https://quizlet.com/923786/bronsted-lowry-theory-of-acids-and-bases-flash-cards/