

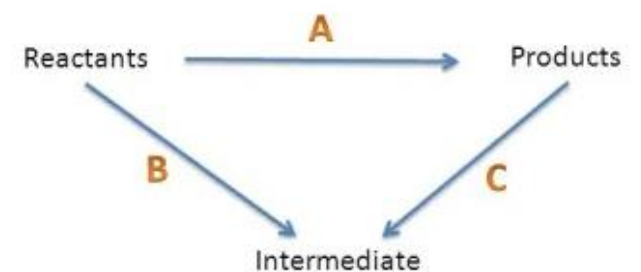
# Hess' law

**Introduction:** Hess' law states that, a reaction can take place using different routes once conditions are the same.

## Hess' Law



Hess' law states that, if a reaction can take place by more than one route and the initial and final conditions are the same, the total enthalpy change is the same.



$$\Delta H_{(\text{Route A})} = \Delta H_{(\text{Route B})} - \Delta H_{(\text{Route C})}$$

Draw an **enthalpy cycle** – then if you follow the direction of an arrow then ADD. If opposite to the direction of the arrow then SUBTRACT.



[https://www.youtube.com/watch?v=chXMpDwjBDk&feature=emb\\_logo](https://www.youtube.com/watch?v=chXMpDwjBDk&feature=emb_logo)

<https://prezi.com/vphfgrbjn9xc/hesss-law/>

<https://quizlet.com/11898091/54-hesss-law-flash-cards/>