

Lattice energy

Introduction: Lattice energy is influenced by two (2) main factors ionic charge and ionic radius.

Lattice energy

- Factors affecting magnitude of lattice energies:
 - a) Ionic charge increase*, more exothermic, *higher* lattice energy.
 - b) Ionic radius decrease*, more exothermic, *higher* lattice energy.
 - c) Crystal structure*

$$\text{Lattice energy} \propto \frac{(q_+ \times q_-)}{(r_+ + r_-)}$$

Compound	Lattice energy / kJmol^{-1}
NaF	-915
NaCl	-776
NaBr	-742
NaI	-699
MgCl_2	-2489
MgO	-3933

https://www.youtube.com/watch?v=Jh78bJEnE_I&feature=emb_logo

<https://prezi.com/vbspaxuzru0/born-haber-cycle/>

<https://quizlet.com/118380970/lattice-energy-flash-cards/>