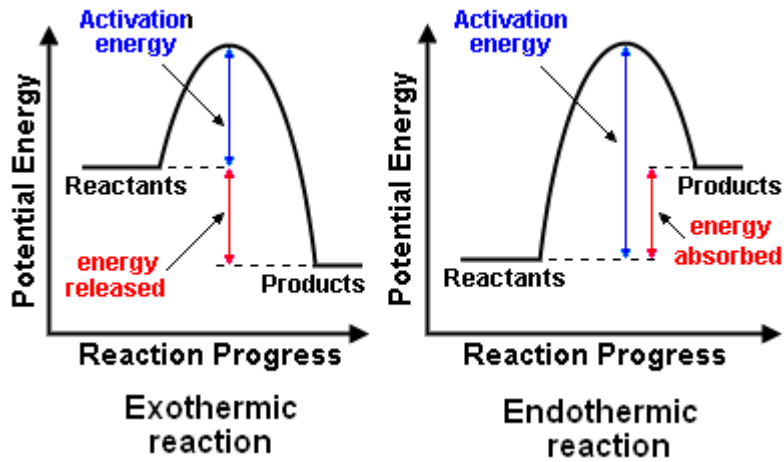


# Energy changes

**Introduction:** Energy changes occur in every reaction. Overall energy is either absorbed (endothermic) or released (exothermic).



Exothermic	Endothermic
<ul style="list-style-type: none"> <li><u>-Releases energy</u></li> <li><u>- Takes a lot of energy to break bonds</u></li> <li><u>- Increase in temperature</u></li> <li><u>- Energy released as heat</u></li> <li><u>- More energy is being released than added.</u></li> </ul>	<ul style="list-style-type: none"> <li><u>-A lot of energy required to break bonds</u></li> <li><u>-Absorbs energy</u></li> <li><u>- More must be added, than released.</u></li> <li><u>- Decrease in temperature</u></li> <li><u>- energy absorbed as electrical energy</u></li> </ul>

<https://prezi.com/wesch0rft0dt/ch-33-energy-changes-in-chemical-reactions/>

<https://quizlet.com/107327872/73-energy-changes-in-reactions-flash-cards/>