

Ideal gas vs. Real gas

Introduction: The properties of ideal gases are different from those of real gases. Real gases however approach ideal behaviour under specific conditions.

Major Differences between ideal and Real gases

II. Ideal vs. Real Gases

| Ideal Gases | Real Gases |
|---|---|
| Follow gas laws of temperature and volume | Do have volume and attraction between particles |
| Conform to kinetic theory of matter | Gases can turn into liquids and solids |
| No volume or attraction to particles within the gas | Differences occur at low temps and high pressures |
| Examples...No gases meet these rules all the time | Look at pg. 429 "Ideal vs. Real Gases" |
| But... | Ex: Anything that is a gas or can become a gas |

Conditions necessary for Real gases to approach Ideal behaviour

https://www.youtube.com/watch?v=5kT7oUB0q6Q&feature=emb_logo

The characteristics of ideal and real gases are compared and the conditions necessary for Real gases to approach Ideal behaviour are explained.

<https://www.slideshare.net/marianjoycemacadine/ideal-gas-law-26204836?ref=https://notesmaster.com/en/group/caribbean/1556-cape-covid19-support/24540-ideal-gas-vs-real-gas>