

Empirical and Molecular formulas

Introduction: Empirical formula gives the simplest whole number ratio of the elements in a compound. The molecular formula is the actual ratio of elements in a compound.

EMPIRICAL FORMULA

V E R S U S

MOLECULAR FORMULA

Empirical formula is the simplest form of expressing the elemental composition of a compound.

The empirical formula is derived first from the weight percentages of the elements present in the compound.

The empirical formula contains the most simplified ratio of the moles of elements in the compound.

The empirical formula is not often used in reaction schemes.

Pediaa.com

Molecular formula is the actual representation of the elemental composition of the compound.

The molecular formula is related to the total weight of the compound in question and often is derived after the obtaining the empirical formula.

The molecular formula needs to be a multiple of the empirical formula.

The molecular formula is commonly used in reactions and other chemical recordings.

https://www.youtube.com/watch?v=wnRaBWvhYKY&feature=emb_logo

https://www.youtube.com/watch?v=AFqwtY7m2PI&feature=emb_logo

<https://prezi.com/4rcgjk2k5gbj/empirical-and-molecular-formulas/>

<https://quizlet.com/11637046/empirical-and-molecular-formula-flash-cards/>